

Errors, Error Propagation, Statistics 510 Lab lecture 02 / 09 / 2009



Why do we quote errors ?



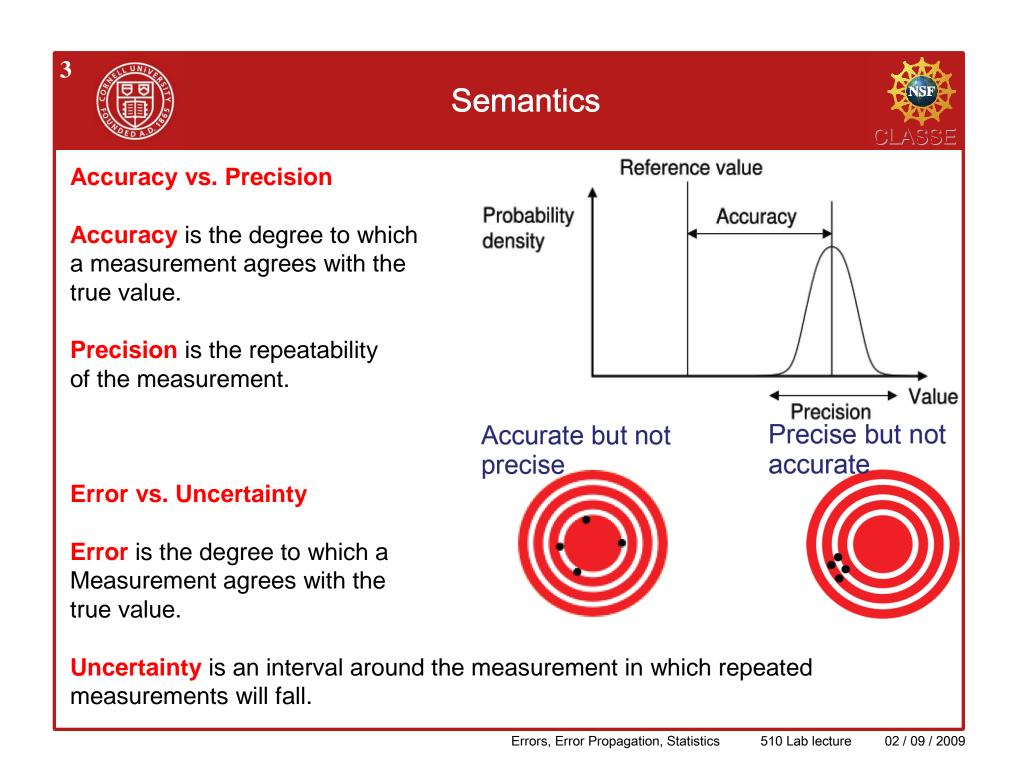
Thanks to Anders Ryd for much of the presented material.

It provides information about the precision of the measurement.

For example the gravitational constant is measured to be $G_N=6.90\ 10^{-11}m^3kg^{-1}s^{-2}$ The 'accepted' value is $G_N=6.6742(10)\ 10^{-11}\ m^3kg^{-1}s^{-2}$.

Without quoting any errors we don't know if this was just a less precise measurement or a Nobel prize worthy discovery.

Example: G_N =(6.90+/-0.25) 10⁻¹¹ m³kg⁻¹s⁻² would be in good agreement G_N =(6.90+/-0.01) 10⁻¹¹ m³kg⁻¹s⁻² would be an interesting result ...





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Error is almost never what we are interested in. In science we typically do not know the 'true' value.

Rather we are interested in the **uncertainty**. This is what we need to quantify in any measurement.

We are often very sloppy and inconsistent in our language and call what is actually an uncertainty an error, **e.g. in the title of this lecture**.

Especially in High Energy Physics we try to get this straight when we write a paper, but in every day talk we are also sloppy and use the word error instead of uncertainty.

When we talk about measurement error,

- It is not a blunder
- It is not an accident
- It is not due to incorrectly handling the equipment
- It is not the difference to an accepted value found in the literatur



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Importance of uncertainty



Example 1:

High fiber diets: A study in 1970 claimed that a high fiber diet reduces polyps forming in the colon, being precursors of cancer. A study in 2000 with more analyzed individuals showed no such effect. The **uncertainty** in the first study was too large an not properly accounted for. This left people eating lots of fibers for 30 years – yuck (just kidding).

Example 2:

A study in the late 60s found large levels of iron, which is required for red blood cell production, in spinach. Popular comics tried to promote spinach consumption.

A study in the 90s showed that the original measurement had a reading error in the decimal point. The iron levels are a factor of 10 lower than claimed. The incorrect reading of the decimal was a **blunder**, not due to an uncertainty in the measurement.



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Fact of scientific life:

Scientists subconsciously bias data to their desired outcome, even when they know about this tendency of their psyche.

Example: N-rays

X-rays discovered in 1895 by Roentgen with huge and fast success.

Another new type of radiation was reported in 1903:

Rene Blondlot (physicist, Nancy / F) discovered N-rays (with N for Nancy)

These became a matter of national pride to the French. Later several scientists, mostly French, claimed to have seen these rays.

100eds of papers published within about one year, 26 from Blondlot.

They go through wood and metal but are blocked by water.

They could be stored in a brick.

They are **emitted by rabbits, frogs and the human brain** (medical imaging) Jean Becquerel (son of Henri who discovered radioactivity) found N-rays **transmitted over a wire** (brains can per telephone ...)

Robert Wood (John Hopkins) went to Blondlot's lab and secretly removed the sample. Blondlot insisted he was still measuring N-rays.

Within month no one believed in N-rays any more.





The **Uncertainty** (Error), in the lab, describes the distance from your measurement result within which your setup has determined that the true value is likely to lie. It describes therefore a property of your measurement procedure, when followed correctly.

You measure N_A in experiment G10 and you obtain the literature value to 10^{-3} . This does not show that your setup has established that the true value of N_A is likely to be within 10^{-3} of your result. The 10^{-3} is therefore not the uncertainty of your experiment.

A new measurement following the same procedure will lead to a different measurement result, but usually the same uncertainty. The new and old result are likely to differ by an amount that is about as large as the uncertainty. The uncertainty is therefore a property of the measurement procedure, and building a good experiment means building an experiment with relatively small uncertainty.



Meaning of an Error



If we measure a voltage V_{sat} = 10.2 +/- 0.3 V, what does this mean?

In general there are differences in different science disciplines.

In physics, a 1sigma error is generally used. If the measurements are normally distributed (Gaussian), this corresponds to a 68% confidence level (CL) interval. Or that 32% of the time the true value would be outside the quoted error range.

For statistical errors, this can be given a precise meaning. May other errors are harder to estimate.



Different types of errors



- Statistical (e.g. labs N2, N4, N15, N17)
 From finite statistics, originates in the Poisson distribution.
- Systematic (All)

e.g. how well can you measure a voltage, length, etc.

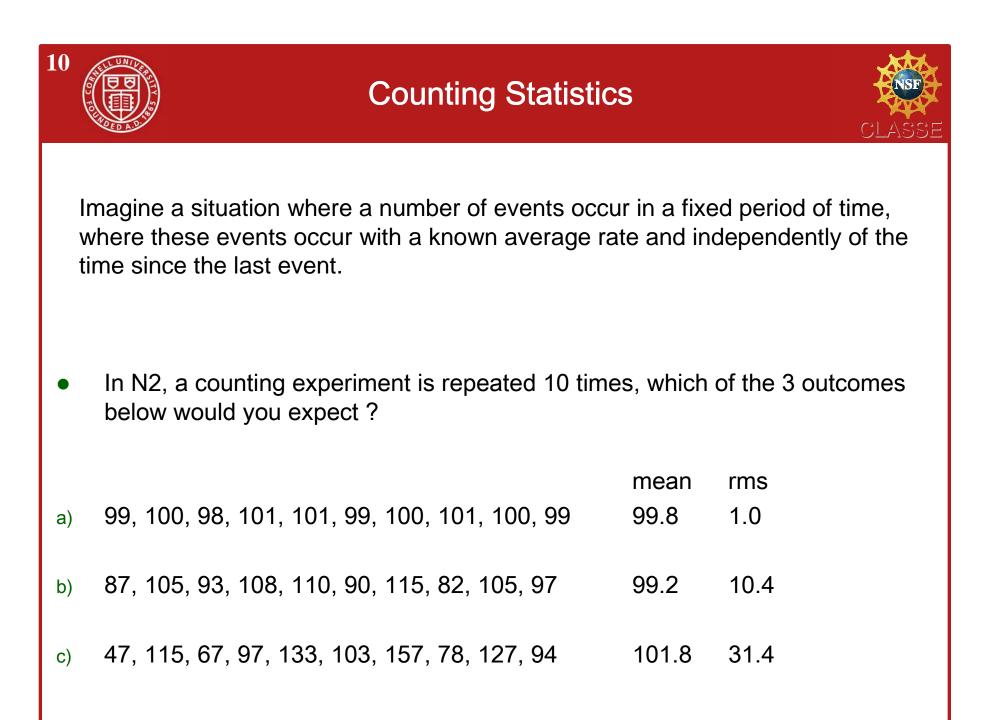
• Theory (N15, N17)

In these experiments you measure the muon lifetime. But there are corrections to the capture rate for mu- that comes from theory.

- It is common to quote these uncertainties separately: $\tau_{\mu} = (2.19 + -0.05_{stat.} + -0.01_{syst.} + -0.02_{th.})\mu s$
- Different notations are used for uncertainties, e.g.

 τ_{μ} =(2.19(5)_{stat.}+/-(1)_{syst.}+/-(2)_{th.}) μ s

N2: Alpha particle range in air and Helium, N4: Ratherford scattering of alpha particles, N15: μ -meson lifetime, N17: μ -meson lifetime with PC

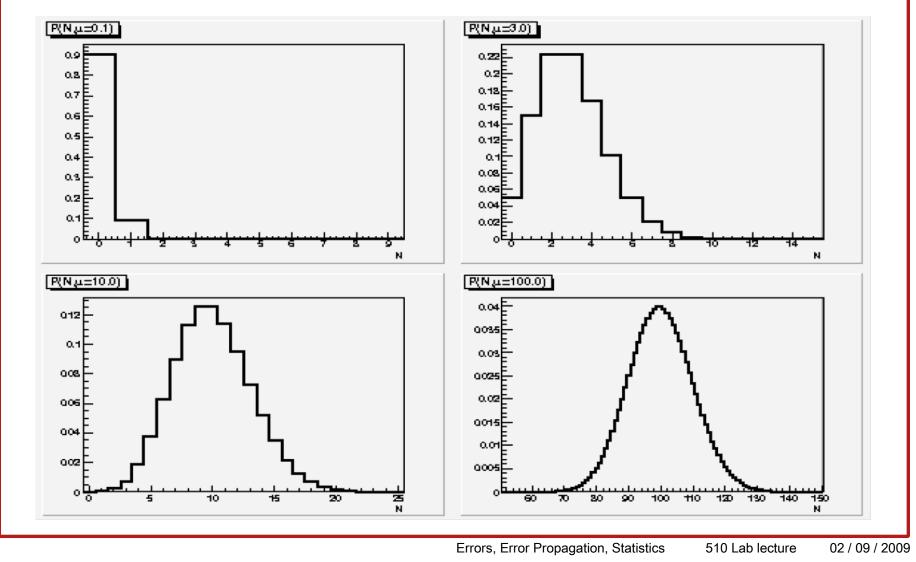




The Poisson Distribution



For large $\langle N \rangle$ ($\mu > 10$), the Poisson distribution approaches a normal distribution.





Poisson Distribution



$$P(N,\mu) = \frac{\mu^N}{N!} e^{-\mu}$$
 $\sum_{N=0}^{\infty} P(N,\mu) = 1$

$$\langle N \rangle = \sum_{N=0}^{\infty} N P(N,\mu) = \sum_{N=1}^{\infty} N \frac{\mu^N}{N!} e^{-\mu} = \mu \sum_{N=1}^{\infty} \frac{\mu^{N-1}}{(N-1)!} e^{-\mu} = \mu$$

$$rms^{2} = \left\langle \left(N - \left\langle N \right\rangle\right)^{2} \right\rangle = \left\langle N^{2} - 2N\left\langle N \right\rangle + \left\langle N \right\rangle^{2} \right\rangle = \left\langle N^{2} \right\rangle - \left\langle N \right\rangle^{2}$$
$$\left\langle N^{2} \right\rangle = \sum_{N=0}^{\infty} N^{2} P(N,\mu)^{2} = \sum_{N=1}^{\infty} N^{2} \frac{\mu^{N}}{N!} e^{-\mu} = \mu \sum_{N=1}^{\infty} N \frac{\mu^{N-1}}{(N-1)!} e^{-\mu}$$
$$= \mu \sum_{N=1}^{\infty} \left[(N-1) \frac{\mu^{N-1}}{(N-1)!} e^{-\mu} + \frac{\mu^{N-1}}{(N-1)!} e^{-\mu} \right] = \mu(\mu+1)$$

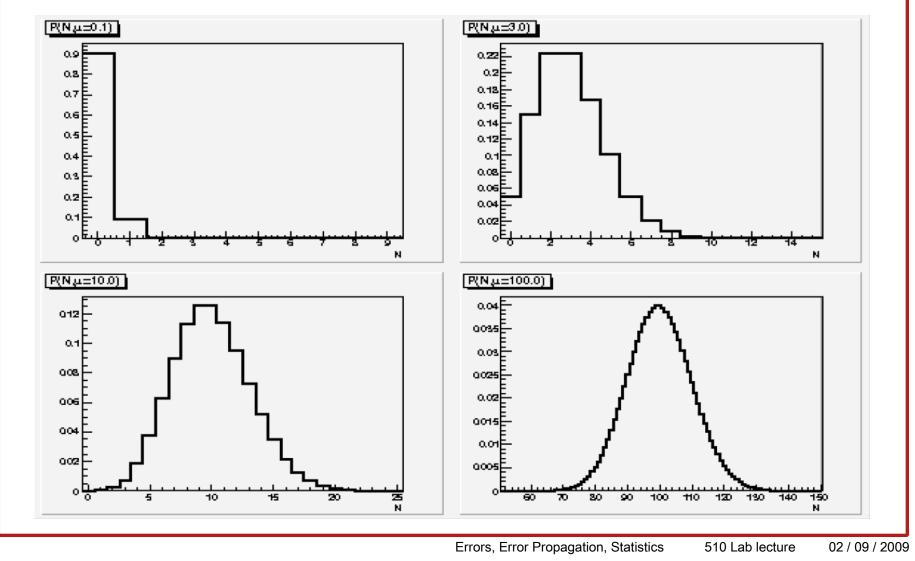
$$rms^{2} = \mu \implies rms = \sqrt{\mu} \implies \frac{rms}{\langle N \rangle} = \frac{1}{\sqrt{\langle N \rangle}}$$



Limit of the Poisson Distribution



For large <N> (μ >10), the Poisson distribution approaches a normal distribution.





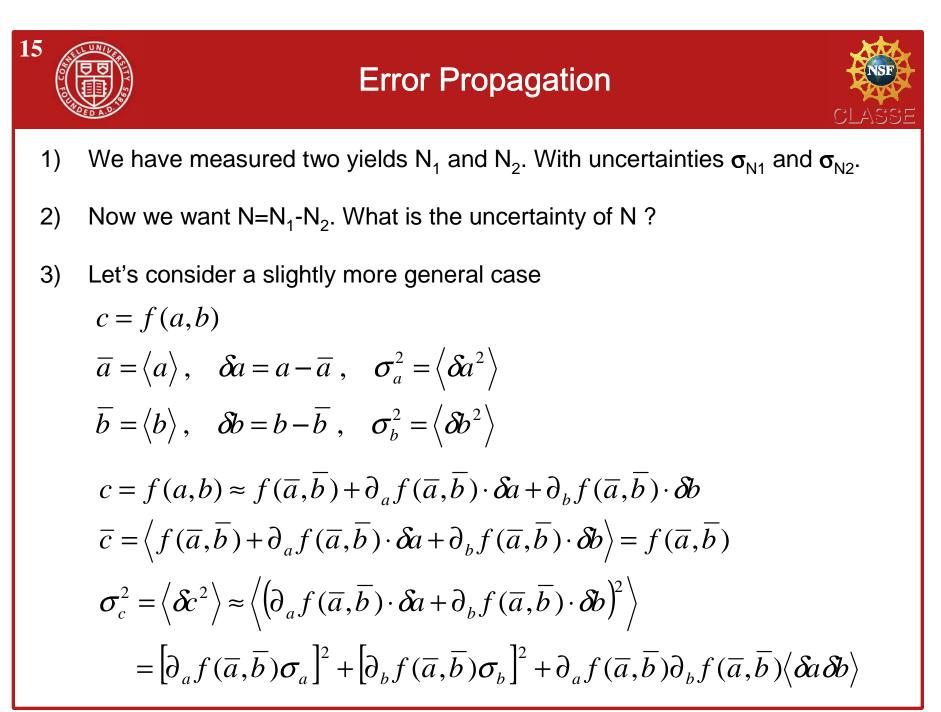
Use of the Poisson Distribution



- The Poisson distribution tells you how probable it is to obtain a given count if the mean is known.
 - Typically we don't know the true mean, but our measured count serves as an estimate of the mean.

We can now use this information to estimate the uncertainty.

E.g. in a counting experiment we obtain 98 counts. We then assign the uncertainty of 98^{1/2} = 9.9 to say that the measurement leads to 98+/-10 counts.





Uncorrelated Errors



Assumption of uncorrelated errors:

Errors in variable a vary independently of those in variable b.

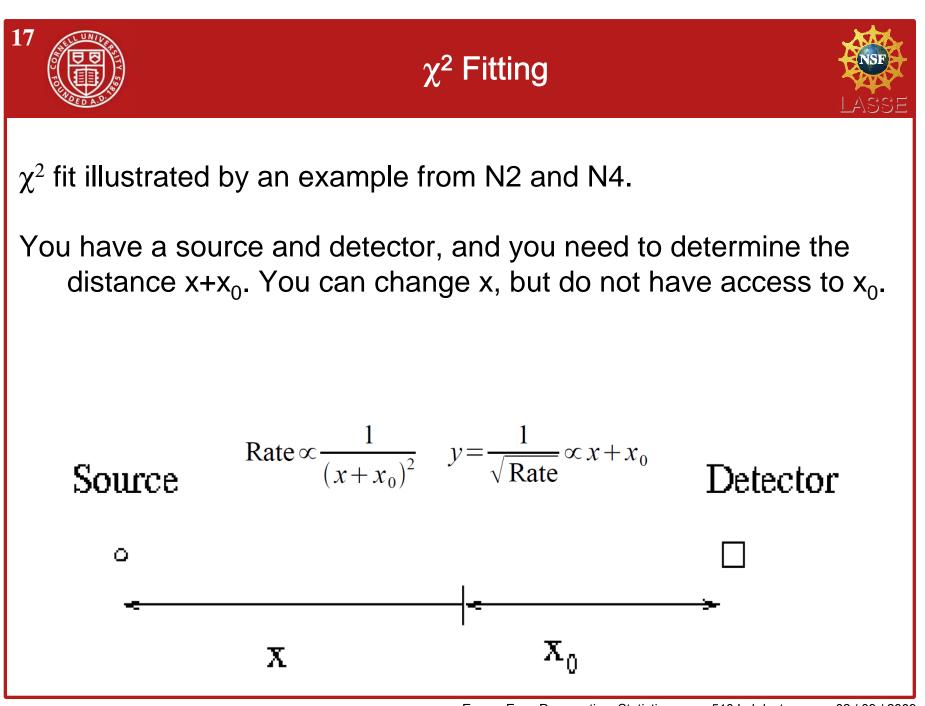
$$\langle \delta a \, \delta b \rangle = \langle \delta b \rangle \langle \delta a \rangle = 0$$

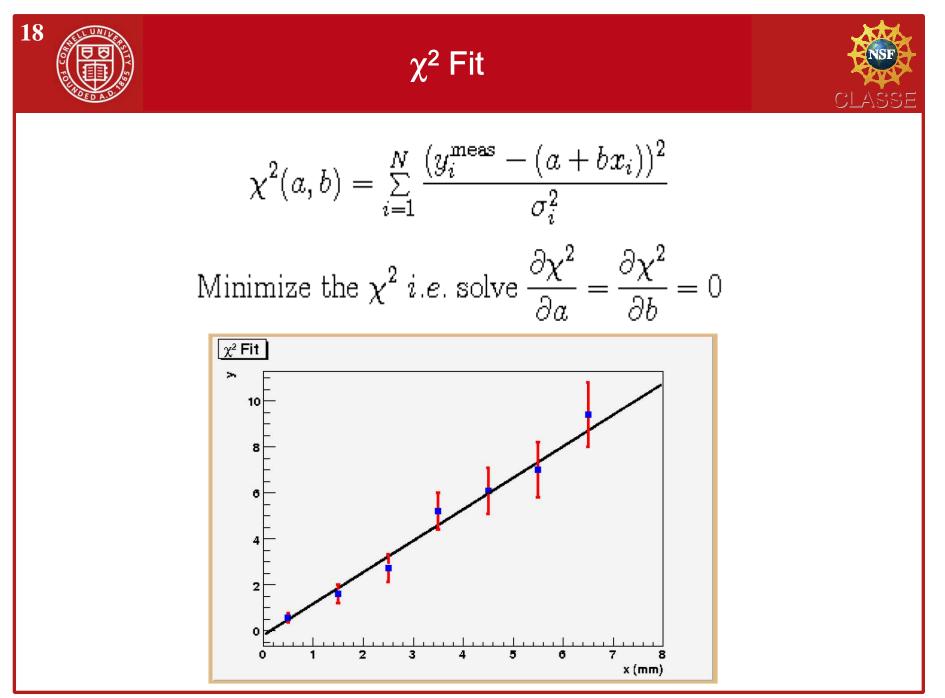
$$\boldsymbol{\sigma}_{c}^{2} = \left[\partial_{a}f(\overline{a},\overline{b})\boldsymbol{\sigma}_{a}\right]^{2} + \left[\partial_{b}f(\overline{a},\overline{b})\boldsymbol{\sigma}_{b}\right]^{2}$$

Example: c = a - b

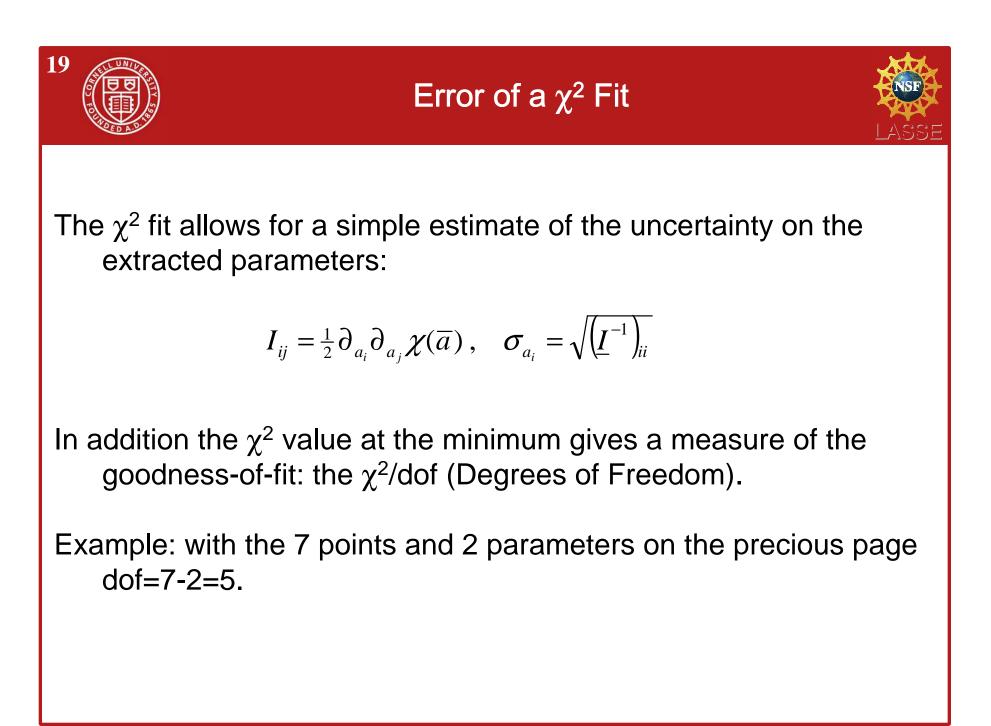
$$\sigma_c^2 = \sigma_a^2 + \sigma_b^2$$

$$\sigma_c = \sqrt{\sigma_a^2 + \sigma_b^2}$$





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If you measured a quantity by M independent procedures and obtained the values c_i with uncertainty σ_i , what is the best combined measurement and uncertainty ?

$$c = \frac{\sum_{i=1}^{M} c_{i}}{\sum_{i=1}^{M} \frac{1}{\sigma_{i}^{2}}}$$

The error propagation formula $\sigma_c^2 = [\partial_a f(\overline{a}, \overline{b})\sigma_a]^2 + [\partial_b f(\overline{a}, \overline{b})\sigma_b]^2$ leads to

$$\sigma_c^2 = \sum_{i=1}^M \left(\frac{\frac{1}{\sigma_i^2}}{\sum_{i=1}^M \frac{1}{\sigma_i^2}} \sigma_i \right) = \sum_{i=1}^M \left(\frac{\frac{1}{\sigma_i}}{\sum_{i=1}^M \frac{1}{\sigma_i^2}} \right) = \frac{1}{\sum_{i=1}^M \frac{1}{\sigma_i^2}}$$

For M identical uncertainties $\sigma_i = \sigma$: $\sigma_c = \frac{\sigma_i}{\sigma_i^2}$

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Resources:

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In the 510 lab: J. Orear L. Hand

Others: http://dcaps.library.cornell.edu/etitles/Frodesen/probabilitystatisticsparticlephysics.pdf Particle Data Group (PDG): http://pdg.lbl.gov/